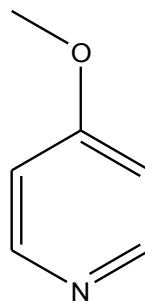
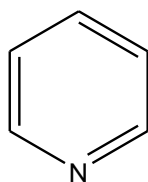
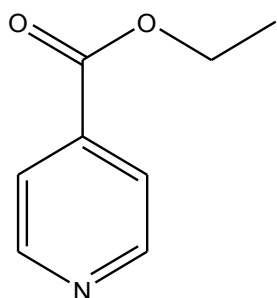


298 K	K_{1S}	2,9	8,2	30,4
	K_2	19,8	22,4	27,2
	K_2/K_{1S}	6,8	2,7	0,9
308 K	K_{1S}	2,16	6	21,5
	K_2	13,9	15,8	20,2
	K_2/K_{1S}	6,5	2,6	0,94
318 K	K_{1S}	1,61	4,5	15,8
	K_2	10,4	11,6	14
	K_2/K_{1S}	6,5	2,57	0,89
328 K	K_{1S}	1,26	3,4	11,6
	K_2	7,65	8,6	10,8
	K_2/K_{1S}	6,07	2,53	0,93

Zusammenfassung der thermodynamischen Konstanten:



K_{1S}	ΔH	-22.81 kJ/mol = -5.45 kcal/mol	-22.13 kJ/mol = -5.29 kcal/mol	-25.93 kJ/mol = -6.19 kcal/mol
	ΔS	-67.68 J/mol = -16.17 cal/mol	-56.98 J/mol = -13.61 cal/mol	-58.6 J/mol = -14.0 cal/mol
K_2	ΔH	-25.49 kJ/mol = -6.09 kcal/mol	-24.999 kJ/mol = -5.97 kcal/mol	-25.457 kJ/mol = -6.08 kcal/mol
	ΔS	-60.78 J/mol = -14.52 cal/mol	-58.1 J/mol = -13.88 cal/mol	-57.87 J/mol = -13.82 cal/mol
β_2	ΔH	-48.30 kJ/mol = -11.54 kcal/mol	-47.09 kJ/mol = -11.25 kcal/mol	-51.188 kJ/mol = -12.23 kcal/mol
	ΔS	-128.46 J/mol = -30.69 cal/mol	-114.98 J/mol = -27.47 cal/mol	-115.82 J/mol = -27.67 cal/mol